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...

Solution: Mole fraction  
of ethanol;  $x = \text{moles}$

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ethanol / (moles ethanol + moles water) = 0.05. Taking 100 moles of solution as the basis the solution is formed by 5 moles of ethanol and 95 moles of water. As molality is the number of moles of solute in 1 kg of solvent we need to calculate the kg of water forming the solution:

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**Concentration terms**  
**| Worksheet**

Solutions Cheat Sheet.  
Solutions Cheat Sheet.  
Solutions are homogeneous mixtures of two or more than two matters. Solutions can be in solid, liquid and gas phases. Solvent is the medium in which matter is dissolved and solute is the matter that dissolves in solvent. For example, in water sugar solution, water is

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the solvent that dissolves sugar and sugar is the solute that is dissolved in water.

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a review of calculations commonly used in the laboratory. Solutions & Solubility Rules: When a solute is dissolved in a solvent we call the mixture a solution.

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concentration and  
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Example, Making  
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Electrolytes and Nonelectrolytes. An

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electrolyte is a compound whose aqueous solution contains ions. When NaCl dissolves in water, the compound dissociates into  $\text{Na}^+$  and  $\text{Cl}^-$  ions. A good test to determine whether or not a compound is an electrolyte is to measure the ability of its water solution to conduct an electrical current.



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# Ions in Solution

(Worksheet)-

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Solutions 1) If I have  
340 mL of a 0.5 M NaBr  
solution, what will the

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concentration be if I  
add 560 mL more  
water to it? 0.19 M (the  
final volume is 900 mL,  
set up the equation  
from that) 2) If I dilute  
250 mL of 0.10 M  
lithium acetate solution  
to a volume of 750 mL,  
what will the  
concentration of this  
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